Chapter 04 Test

Name:_____

Date:_____

1 Which is a property shared by most molecular compounds?

A high boiling point
B high melting point
C low melting point
D nonpolar bonds

2 When an atom loses an electron, it becomes a

A positive ion.
B negative ion.
C neutral ion.
D neutral atom.

3 In the chemical formula for an ionic compound, which item is written first?

A the name of the positive ion
B the name of the negative ion
C the subscripts
D the charges

4 The chemical name for the compound with the formula Na_2S is

A sodium fluoride.
B magnesium sulfide.
C lithium oxide.
D sodium sulfide.

5 What is a double bond?

A a bond between two atoms
B one pair of electrons shared between two atoms
C two pairs of electrons shared between two atoms
D two pairs of electrons shared between four atoms

6 Electrons involved in bonding between atoms are

A valence electrons.
B inside the nucleus.
C closest to the nucleus.
D positively charged.

7 A mixture made of two or more elements, at least one of which is a metal is called a(n)

A superconductor.
B metallic bond.
C complex metal.
D alloy.

8 Which of the following terms means that metals can be rolled into thin sheets, as in aluminum foil, or beaten into complex shapes?

A polar
B alloy
C ductile
D malleable

9 Which of the following groups contains the most reactive elements?

A the alkali metals
B the alkaline earth metals
C the carbon family
D the noble gases

10 Which of the following is a property of ionic compounds?

A They have low melting points.
B They have low boiling points.
C They form hard, brittle crystals.
D They contain no charged particles.

11 Water is polar and oil is nonpolar. What happens when the two liquids are poured into the same container?

A Both liquids become nonpolar.
B A gas is produced.
C The liquids mix well.
D The liquids do not mix.

12 Which of the following best describes a metal solid?

A metal atoms held together by covalent bonds
B metal atoms held together by ionic bonds
C positive metal ions surrounded by freely moving valence electrons
D neutral metal atoms surrounded by freely moving valence electrons